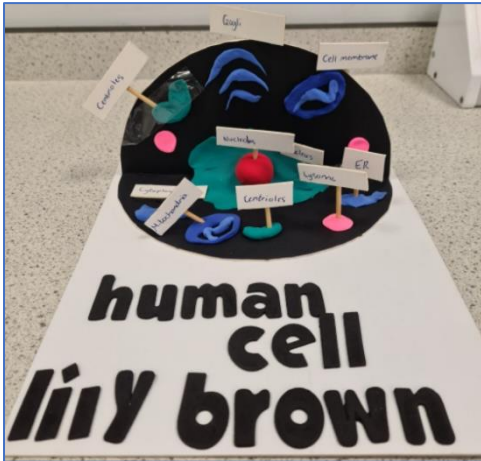
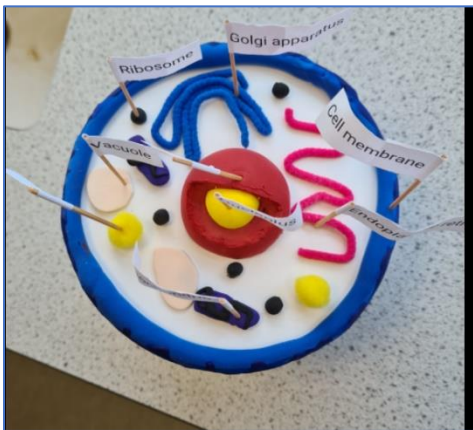


# Work of the week – 22<sup>nd</sup> November 2021

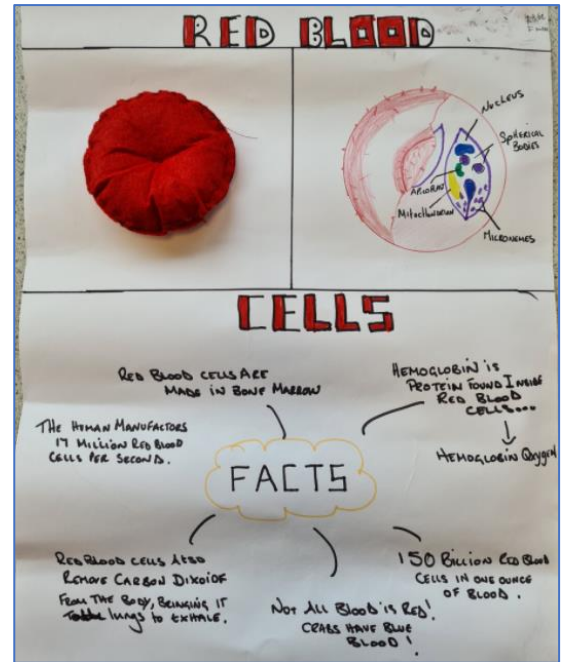


Year 7  
Science  
Curiosity shown in this science homework to make a model of a cell.



Year 7  
Science  
Using fabric to represent a red blood cell and including lots of key facts about the cell. Excellent.

Year 7  
Science  
A 3D model of a cell with all the organelles clearly labelled. Great curiosity.



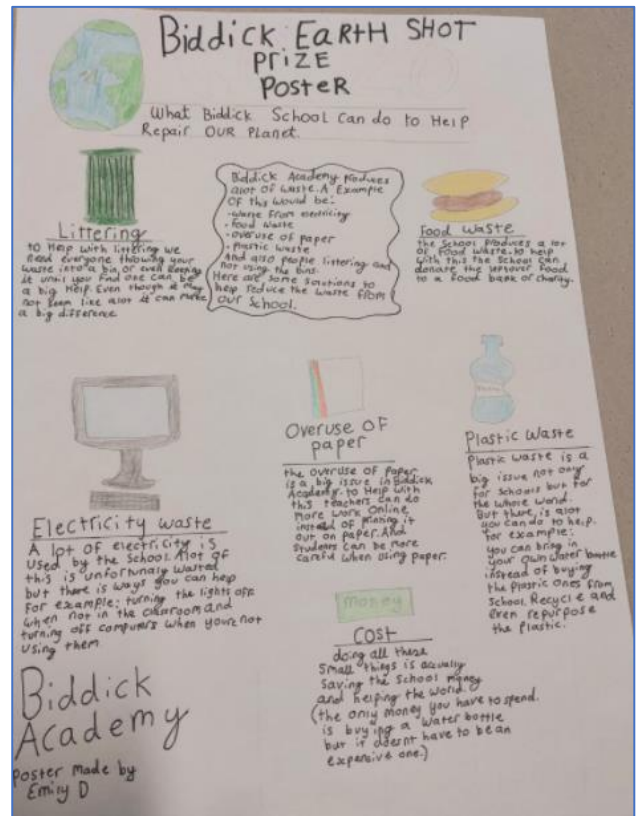
Compare and contrast

Group 1 are very reactive when in water, an example is potassium. When potassium comes in contact with water it produces a spark like flame, which is purple/blue. This depends on how much potassium is used. Group 1 elements are all alkali metals, these elements are so soft that you can cut them with a butter knife. Moreover, Group 1 elements are low density so they will float in water. Also Group 1 are good conductors of electricity. The most reactive element in Group 1 is Francium and the least reactive is Lithium. I know this because the further you go down the more reactive they become.

Group 7 are all toxic and toxic/harmful, an example is Fluorine. (Fluorine is the most reactive element gas in Group 7) The least reactive in Group 7 is Iodine, however this does not mean it is not toxic/harmful in any way. Group 7 is also known as Halogens and most elements in group 7 kills bacteria. Moving onto Group 0/8, Group 0 is also known to as Noble Gases. They are all gases along with Group 7. The elements in Group 0 are not very reactive, for example Helium, Helium is used in balloons and it is not harmful. Another example: Neon - is used in neon signs. Helium is also used when going diving (it is in the oxygen tanks).

*Wow - This is phenomenal work - you have given a great summary of the 3 different groups!*

Year 8 Science  
Brilliant comparison of three groups in the periodic table showing aspiration.



Year 10  
An informative poster to make us aware of how changing little things can have a positive impact on the environment. Well done.